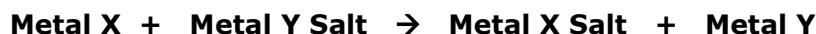


# Displacement of Metals

# Chemical 16 Equations

## General Rule



A metal will be displaced from a salt by a **more reactive** metal.

e.g. 1. Magnesium + Lead Oxide  $\rightarrow$  Magnesium Oxide + Lead

A metal will not be displaced from a salt by a **less reactive** metal.

e.g. 2. Lead + Magnesium Oxide  $\rightarrow$  No Reaction

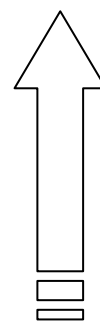
### Task 1

**In your exercise book**, write word equations for the following reactions. If you have been taught how to, then write a balanced symbol equation under each word equation (assume a valency of 3 for iron, 2 for lead and 1 for copper).

	Reactants
a.	Aluminium and lead chloride
b.	Potassium and calcium chloride
c.	Sodium and iron chloride
d.	Zinc and aluminium chloride
e.	Lead and copper chloride
f.	Iron and copper oxide
g.	Calcium and lead iodide
h.	Tin and zinc sulfate
i.	Calcium and copper sulfate
j.	Sodium and potassium fluoride
k.	Magnesium and lead nitrate
l.	Iron and zinc sulfate
m.	Copper and aluminium carbonate
n.	Potassium and magnesium chloride
o.	Zinc and lead sulfate
p.	Magnesium and calcium iodide
q.	Aluminium and lead bromide
r.	Potassium and sodium carbonate
s.	Calcium and copper nitrate
t.	Zinc and copper sulfate

### Reactivity Series

most  
reactive



least  
reactive

Potassium  
Sodium  
Calcium  
Magnesium  
Aluminium  
Zinc  
Iron  
Tin  
Lead  
Copper